26 Diborane has the formula B_2H_6.

Assume that boron consists of two isotopes, containing 20% $^{10}_{5}B$ atoms and 80% $^{11}_{5}B$ atoms, and that all hydrogen atoms are $^{1}_{1}H$.

Molecules of diborane will therefore have relative masses of 26, 27 or 28.

In what relative proportion will molecules of diborane with masses of 26, 27 and 28 occur?

A 1 : 2 : 8

B 1 : 2 : 16

C 1 : 4 : 8

D 1 : 4 : 16

E 1 : 8 : 16

F 1 : 8 : 64

G 1 : 16 : 64